

This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1921 Excerpt: ...levo tartaric acids, is due to external compensation. The dextro rotation of the dextro acid is neutralized by the levo rotation of the levo acid. Since racemic acid and the tartaric acids were the first optically active substances to be investigated, the term racemic is often used as a general expression to designate an inactive substance which may be decomposed into its optically active constituents. Levo tartaric acid may be prepared from racemic acid by the general methods we have just described. It melts at 1700 and resembles ordinary dextro tartaric in its general chemical behavior, except in its action on polarized light. It rotates the plane of polarization to the left, while dextro tartaric rotates to the right. The number of degrees of rotation is the same in each case. The sodium ammonium salts of the dextro and levo acids, as we have previously explained, crystallize in forms which are the mirror images of each other. Mesotartaric acid is obtained together with racemic acid when dextro tartaric is heated with water at 165 for about 2 days. It is also formed along with racemic acid by the interaction of dibrom succinic acid and silver hydroxide. Mesotartaric acid is rather more soluble than racemic acid. The acid potassium salt of mesotartaric acid is also more soluble than the acid potassium salt of dextro or levo tartaric acids. Mesotartaric acid crystallizes with one molecule of water of crystallization, and the anhydrous acid melts at 1400. Its solutions are optically inactive and it cannot be decomposed into optically active isomers. Stereochemistry of the tartaric acids. As we have previously explained, there are four dihydroxy succinic acids having the is crystalline and contains two asymmetric carbon atoms. Dextro and levo t...

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